soaps of 12 or more C atoms, shows the same effect as the Na and Ag soaps. (*Chem. Abs. 42,* 6198.)

SOLUBILIZATION OF DIMETHYLAMINOBENZENE IN SOLUTIONS OF DETERGENTS. I. THE EFFECT OF TEMPERATURE ON THE SOLUBILIZATION AND UPON THE CRITICAL CONCENTRATION. I. M. Kolthoff and W. Stricks. *J. Phys. & Colloid Chem. 52,* 915-41(1948). Solubility of a dye occurs when the detergent concentration reaches a critical value. Above this value, the solubility of the dye is proportional to the detergent concentration. The critical concentration is independent of temperature between 30° and 50°. With fatty acid soaps, the critical concentration decreases rapidly with increasing carbon number of the fatty acid. Solubilizing power of different detergents increases with increasing temperature. (*Chem. Abs. 42,* 7070.)

OLEATE SYSTEMS CONTAINING POTASSIUM CHLORIDE IN WHICH THE POTASSIUM CHLORIDE CONCENTRATION IS STILL TOO LOW FOR COACERVATION. I. THE VISCOSITY-POTASSIUM CHLORIDE CURVE. H. G. Bungenberg de Jong and G. W. H. M. van Alphen (Univ. of Leiden, Netherlands). *Proc. Koninkl. Nederland. Akad. Wetenschap. 50,* 849-57(1947). (In English). By use of an Ostwald viscometer the effect of KCl on dilute aqueous solutions of Na oleate (I) was investigated at 25° in a preliminary manner. Over the concentration range of 0.057 to 1.36% I the viscosity curve, as a function of increasing concentration of KCl, rises to a maximum, then decreases sharply before coacervation occurs. For all these concentrations of I coacervation occurs at about the same concentration of KCl (9 ml. 3.8 N KCl/20 ml.). The ascending branch of the curve contains a point of reflection, at which a "thread-pulling" property is first exhibited. Both the maximum and the inflection point are depressed by increasing the temperature (36, 49°). The systems are very sensitive to contamination. (*Chem. Abs. 42,* 6609.)

CRITICAL EFFECT OF TEMPERATURE ON THE ABSORPTION OF WATER BY SOLUTIONS OF ETHANOLAMINE OLEATE IN BENZENE. R. C. Pink (Queen's Univ. Belfast, Ireland). *Trans. Faraday Soc. 42B,* 170-3(1946). The amount of water absorbed by benzene solutions of ethanolamine oleate increases greatly above a critical temperature of 50°. This change is due to a change in state of the micelle. Above the critical temperature liquefaction of the soap occurs, permitting expansion of the swollen micelles to droplets of visible dimensions. (*Chem. Abs. 42,* 6610.)

THE OSMOTIC ACTIVITY AND CONDUCTIVITY OF AQUEOUS SOLUTIONS OF SOME TYPICAL COLLOIDAL ELECTROLYTES. Ann Cushman, A. P. Brady, and J. W. McBain (Stanford U., Calif.). *J. Colloid Sci. 3,* 425-36 (1948). The osmostic coefficients of aqueous solutions of different anionic and cationic colloidal electrolytes have been measured by freezing point depression. Conductivities were also measured. The relative lowering of conductivity with increasing concentration is less than that of osmotic coefficient. Addition of dimethyl phthalate, a substance insoluble in oil and water but solubilized by soap solutions, appears slightly to decrease colloid formation as reflected in osmotic coefficient and equivalent.

MIXED MONOLAYERS. I. ADSORBED FILMS AT AIR-WATER SURFACES. Eric Hutchinson (Stanford Univ., Calif.). *J. Colloid Sci. 3,* 413-24(1948). Measurements have been made of the surface tensions of solutions containing octyl alcohol and sodium dodecyl sulfate.

CORRECTION

Annual Review of Literature on Fats, Oils and Soap. M. M. Piskur (April 1948).

Page 129, right hand column, line 5-6: change *Manuf. Chemist* 28 to *Manuf. Chemist* 18.

Page 133, left hand column, line 14: change *Nor. 70,291-2* to *Nor. 70,292.*